. CHAPTER 3 MATTER

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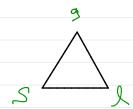
CHAPTER 3 MATTER

Matter

→ has mass

→ occupies space

3 States of Matter



	Constant Shape?	Constant Volume?	Compres sible?
G	X	X	
L	X		X
S			X

Properties: Chemical and Physical

CHEMICAL CHANGE — a change in the chemical identity of a substance (i.e., a "reaction" occurs to change the substance from on thing into another CHEMICAL PROPERTY — an aspect of a substance only observable when there is a chemical change ("reaction")

PHYSICAL CHANGE — a change in a substance in which the chemical identity does NOT change. PHYSICAL PROPERTY — an aspect of a substance that can be observed without changing the chemical identity of a substance.

General Rule of Change:

If you (chemically) end up with what you started with, then it's a Physical Change

^{*} only exception is the property of "inertness", which is considered a Chemical property.

(EX) ¿ID Type of Property? Physical only, Chemical only, or Both?

P · Ice melts at °C

• Propane burns to CO₂ + H₂O

• Pb is a dense metal

• Milk turns sour

• Leaves turn color in Fall

• Gallium melts in your hand

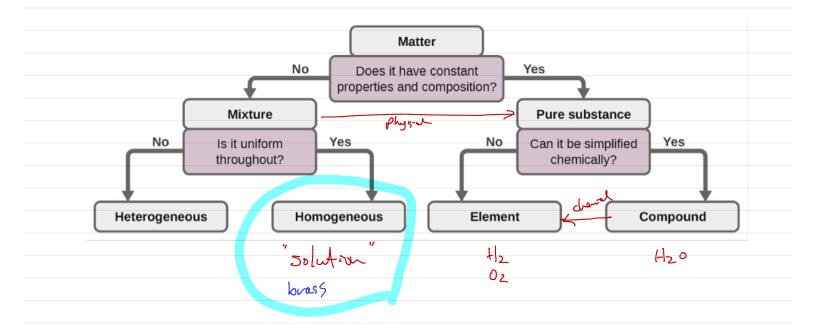
• Scramble an egg

• A nut rusts onto a bolt

• Strike a match

• Au is inert (is not very reactive)

4 Classes of Matter



(EX) ¿ID the class of matter?

- homogeneous mix solution

 NaCl compound

 heterogeneous mixture

 Cu & Zn do not react; merely mixed solution

 complex blend of HC's homogeneous mixture

 homogeneous mix solution

 two phases: oil & water heterogeneous mixture
- Pure apple juice
- Table salt
- Chocolate chip cookie
- Brass (Cu + Zn)
- Gasoline
- Maple syrup
- Oil/vinegar dressing

Separation of Mixtures

- folfretur — A size - dotalet — A B.P. - certification — A density

 \hookrightarrow filtration

 \hookrightarrow distillation